

Egyenletek III.

Osztályzat: «grade»

Tárgy: Kémia

Dátum: «date»

1 ammónia és víz reakciója

- A $\text{NH}_3 + \text{H}_2\text{O} = \text{NH}_4^+ + \text{OH}^-$
- B $\text{NH}_3 + \text{H}_2\text{O} = \text{NH}_3^+ + \text{OH}^-$
- C $\text{NH}_4 + \text{H}_2\text{O} = \text{NH}_4^+ + \text{OH}^-$
- D $\text{NH}_3 + \text{H}_2\text{O} = \text{NH}_4^+ + \text{OH}^+$

2 ammónia és sósav reakciója

- A $\text{NH}_3 + \text{HCl} = \text{NH}_3^+ + \text{Cl}^-$
- B $\text{NH}_3 + \text{NaCl} = \text{NH}_4^+ + \text{Cl}^-$
- C $\text{NH}_3 + \text{HCl} = \text{NH}_4^+ + \text{Cl}^-$
- D $\text{NH}_3 + \text{HCl} = \text{NH}_4^- + \text{Cl}^+$

3 ammónia előállítása

- A $\text{N}_2 + 3\text{H}_2 = 2\text{NH}_3$
- B $\text{N} + 3\text{H} = \text{NH}_3$
- C $\text{N}_2 + 3\text{H}_2 = \text{NH}_3$
- D $2\text{N} + 3\text{H}_2 = 2\text{NH}_3$

4 salétromsav oldódása vízben

- A $\text{HNO}_3 + \text{H}_2\text{O} = \text{NO}_3^- + \text{H}_3\text{O}^+$
- B $\text{HNO}_3 + \text{H}_2\text{O} = \text{NO}_3^+ + \text{H}_3\text{O}^-$
- C $\text{HNO}_3 + \text{H}_2\text{O} = \text{NO}_4^- + \text{H}_3\text{O}^+$
- D $\text{NH}_3 + \text{H}_2\text{O} = \text{NO}_3^- + \text{H}_3\text{O}^+$

5 az ammónia és a salétromsav reakciója

- A $\text{NH}_4 + \text{HNO}_3 = \text{NH}_4^+ + \text{NO}_3^-$
- B $\text{NH}_3 + \text{HNO}_4 = \text{NH}_4^+ + \text{NO}_3^-$
- C $\text{NH}_3 + \text{HNO}_3 = \text{NH}_4^+ + \text{NO}_3^-$
- D $\text{NH}_3 + \text{HNO}_3 = \text{NH}_3^+ + \text{NO}_3^-$

6 a salétomsav és a nátrium-hidroxid reakciója

- A $\text{HNO}_3 + \text{NaOH} = \text{NaNO}_4 + \text{H}_2\text{O}$
- B $\text{HNO}_3 + \text{NaOH} = \text{NaNO}_3 + \text{H}_2\text{O}$
- C $\text{HNO}_3 + \text{NaOH} = \text{NaNO}_3 + \text{H}_2$
- D $\text{HNO}_3 + \text{NaOH} = \text{NaNO}_3 + 2\text{H}_2\text{O}$

7 a szén égése

- A $\text{C} + \text{O}_2 = \text{CO}_2$
- B $\text{C} + \text{O}_2 = \text{CO}$
- C $\text{C} + \text{O} = \text{CO}_2$
- D $\text{C} + \text{O} = \text{CO}$

8 a szén redukálja a vasoxidot

- A $\text{Fe}_2\text{O}_3 + 3\text{C} = 2\text{Fe} + 3\text{CO}_2$
- B $\text{Fe}_2\text{O}_3 + 3\text{C} = 3\text{Fe} + 2\text{CO}$
- C $\text{Fe}_2\text{O}_3 + 3\text{C} = 2\text{Fe} + 3\text{CO}$
- D $\text{Fe}_2\text{O}_3 + 3\text{C} = 2\text{Fe} + 2\text{CO}_2$

9 a szén-monoxid égése

- A $\text{CO} + \text{O} = \text{CO}_2$
- B $2\text{CO} + \text{O} = 2\text{CO}_2$
- C $2\text{CO} + \text{O}_2 = \text{CO}_2$
- D $2\text{CO} + \text{O}_2 = 2\text{CO}_2$

10 a szén redukálja a szén-dioxidot

- A $\text{CO}_2 + \text{C} = \text{CO}$
- B $\text{CO}_2 + \text{C} = 2\text{CO}$
- C $\text{CO}_2 + \text{C} = \text{CO}_2$
- D $\text{CO}_2 + 2\text{C} = 2\text{CO}$

11 a vasoxidot redukálja a szén-monoxid

- A $\text{Fe}_2\text{O}_3 + 3\text{CO} = 2\text{Fe} + 3\text{CO}_2$
- B $\text{Fe}_2\text{O}_3 + 3\text{CO}_2 = 2\text{Fe} + 3\text{CO}_2$
- C $\text{Fe}_2\text{O}_3 + 2\text{CO} = 2\text{Fe} + 2\text{CO}_2$
- D $\text{Fe}_2\text{O}_3 + \text{CO} = 2\text{Fe} + \text{CO}_2$

12 a szén-dioxid oldódása vízben

- A $2\text{CO}_2 + 2\text{H}_2\text{O} = 2\text{H}_2\text{CO}_3$
- B $\text{CO}_2 + \text{H}_2\text{O} = 2\text{H}_2\text{CO}_3$
- C $\text{CO}_2 + \text{H}_2\text{O} = \text{H}_2\text{CO}_2$
- D $\text{CO}_2 + \text{H}_2\text{O} = \text{H}_2\text{CO}_3$

13 a meszes víz és a szénsav reakciója

- A $\text{Ca}(\text{OH})_2 + \text{H}_2\text{CO}_3 = \text{CaCO}_3 + 2\text{H}_2\text{O}$
- B $\text{Ca}(\text{OH})_2 + \text{H}_2\text{CO}_3 = \text{CaCO}_3 + \text{H}_2\text{O}$
- C $\text{Ca}(\text{OH})_2 + \text{H}_2\text{CO}_3 = 2\text{CaCO}_3 + 2\text{H}_2\text{O}$
- D $\text{Ca}(\text{OH})_2 + 2\text{H}_2\text{CO}_3 = \text{CaCO}_3 + 2\text{H}_2\text{O}$