

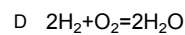
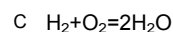
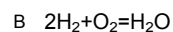
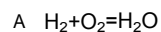
Egyenletek I.

Osztályzat: «grade»

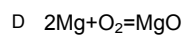
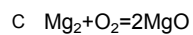
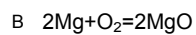
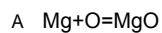
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Dátum: «date»

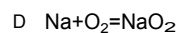
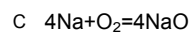
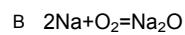
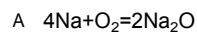
1 a hidrogén égése



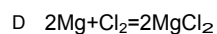
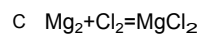
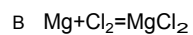
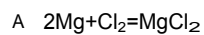
2 a magnézium égése



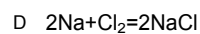
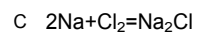
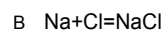
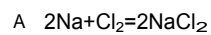
3 a nátrium égése



4 a magnézium és a klór reakciója



5 a nátrium és a klór reakciója



6 a szén égése

- A $C+O_2=CO_2$
- B $2C+O_2=2CO_2$
- C $C+O_2=2CO_2$
- D $C+O_2=2CO$

7 a kén égése

- A $S+O_2=SO_2$
- B $S_2+O_2=S_2O_2$
- C $2S+O_2=S_2O$
- D $2S+O_2=SO_2$

8 a nátrium ionná alakulása

- A $Na=Na^++e^+$
- B $Na=Na^{2+}+2e^-$
- C $Na=Na^-+e^-$
- D $Na=Na^++e^-$

9 a klór ionná alakulása

- A $Cl+e^-=Cl^+$
- B $Cl+2e^-=Cl^{2-}$
- C $Cl+e^-=Cl^-$
- D $Cl+2e^-=Cl^{2+}$

10 a magnézium ionná alakulása

- A $Mg=Mg^{2+}+e^-$
- B $Mg=Mg^++2e^-$
- C $Mg=Mg^{2+}+2e^-$
- D $Mg=Mg^++e^-$

11 az oxigén ionná alakulása

- A $O+2e^-=O^{2-}$
- B $O+e^-=O^-$
- C $O+2e^-=O^{2+}$
- D $O+e^-=O^{2-}$

12 a hidrogén és a klór reakciója

